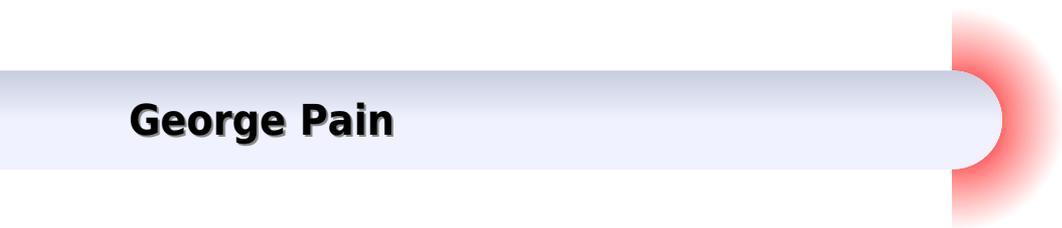


# YOUTUBE AUTOMATION



# Parents Step By Step Guide To YouTube Automation Channel Cheap Starter

**George Pain**



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calculated mass of the product expected x100 in this reaction a yield 80 is anticipated

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 $0 \text{ g Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  c determine the percent yield  
yield actual x 100  $260.0 \text{ g} \times 100 = 86.01$  theoretical 302.3 g

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c is the limiting reagent 4.33 g of h 2 are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s  
remains when the reaction is complete

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yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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 $\text{CH}_3\text{OH}$  can be produced by the following reaction  $2 \text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of co is  
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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  x 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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