

ARBITRATION

ESSENTIAL CONCEPTS

STEVEN C. BENNETT, ESQ.

Arbitration Essential Concepts

David Kirk



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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 & 8 are similar to problem 5

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web percent yield is a measure of the efficiency of the experimental design yield efficiency $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $\frac{160 \text{ g } \text{Al}}{27 \text{ g } \text{Al}} \times 1 \text{ mol } \text{Al} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mol} = 302.3 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7Na₂C₂O₄ is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h₂ are left over 13

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