

Chapter 3 Study Guide – Functional Properties

- 1) Polymeric systems (plastics) are the majority of the bulk of the objects (products) in the 21st century. Unlike all the materials, or even the majority of them, in families with the greatest diversity: **thermo plastics**, **thermosetting plastics**, **elastomers** and **biodegradable plastics**.
- 2) Describe how a polymer bond (essentially different molecules bound together) differs from metallic bonds.
- 3) Describe with the processing of most polymers why these properties are important (10).
- 4) Discuss the relationship between crystallinity in polymers, stiffness, strength, toughness, and other properties of the material.
- 5) Discuss the relationship between crystallinity, toughness, and elongation at break with respect to both natural materials (e.g. wood) and synthetic materials (e.g. engineering polymers) across all the major groups of biological materials, natural fibers, sports plastics, and natural rubbers.
- 6) Describe with the characteristics of organic compounds (polymers), for example, **amorphous**, **elastomers**, **thermoset** groups, etc.
- 7) Describe to the 4 major groups of biological materials (cellulose, proteins, lipids, and water) which are able to:
 - a) Describe the relationships and basic characteristics regarding the structure and polymer of each.
 - b) How are the relationships among and between each
 - c) Characteristic elements

A Study Guide To Chemical Principles 2nd Ed

Mary Burnham, Carol Hurd



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Student Study Guide, Introduction to Chemical Principles, Second Edition Owen C. Gayley, 1986

Study Guide for General Chemistry, Principles and Structure, 2nd Ed., SI Version James E. Brady, Henry Heikkinen, 1980

Catalog of Copyright Entries. Third Series Library of Congress. Copyright Office, 1976

Study Guide and Solutions Manual for Organic Chemistry Joel Karty, Marie M. Melzer, 2014 Written by two dedicated teachers this guide provides students with fully worked solutions to all unworked problems in the text Every solution follows the Think Solve format used in the textbook so the approach to problem solving is modeled consistently The Think step trains students to ask the right questions as they approach a problem and the Solve step then walks them through the solution

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Journal of Chemical Education, 1929

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