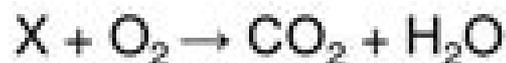


Combustion Analysis



Find out if there's oxygen or unknown compounds in X.

Subtract the mass of C and H from the original mass

$$0.250 \text{ g of compound X} - (0.187 \text{ g C} + 0.063 \text{ g H}) = \underline{0 \text{ g}}$$

So compound X must contain only C and H only and NO oxygen !!

Step 2. Find the number of moles of C and H

$$\text{C: } 0.187 \text{ g} \times (1 \text{ mole C} / 12.01 \text{ g of C}) = 0.0156 \text{ moles C}$$

$$\text{H: } 0.063 \text{ g} \times (1 \text{ mole H} / 1.008 \text{ g of H}) = 0.063 \text{ moles H}$$

Combustion Emission And Analysis P162

M Tight



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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Combustion Emission And Analysis P162**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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