



# Defense Industry

**Jiyuan Zhang**



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subsequent  $\Delta h$  values  $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \frac{5}{2} \text{O}_2 \text{ (g)}$   $\Delta h = 2 \text{ H}_2 \text{ (g)} + \text{C}_2\text{H}_6 \text{ (g)} \rightarrow \Delta h = 94.5 \text{ kJ}$   $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} + \frac{1}{2} \text{O}_2 \text{ (g)} \rightarrow \Delta h = 71.2 \text{ kJ}$   $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{O}_2 \text{ (g)} \rightarrow$

co<sub>2</sub> g 3 h<sub>2</sub> o g Δh 283 kj g 1 c<sub>2</sub>h<sub>6</sub> g

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web hess s law labs by austin lee alayna baron lily zmachinski introduction in order to calculate the enthalpy change for the combustion of magnesium oxide mg s 1 2o<sub>2</sub>g mgo s we used

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web thus for reaction 9 2 in which 1 mole of  $h_3o^+$  and 1 mole of  $oh^-$  combine to form 2 moles of  $h_2o$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{products} - \Delta h_{reactants}$  is positive if heat is absorbed that is if  $\Delta h_{products}$

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