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so calculating enthalpy change of reaction in 2 $\Delta H_{\text{reaction}} = 6 \times \Delta H_{\text{f}}^{\circ}(\text{C}_2\text{H}_6) - 6 \times \Delta H_{\text{f}}^{\circ}(\text{H}_2\text{O}) - 6 \times \Delta H_{\text{f}}^{\circ}(\text{O}_2)$ $2 \times 413 = 6 \times 523 - 6 \times 217 - 6 \times 0$ $826 = 3140 - 1302 - 0$ $826 = 1838$

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