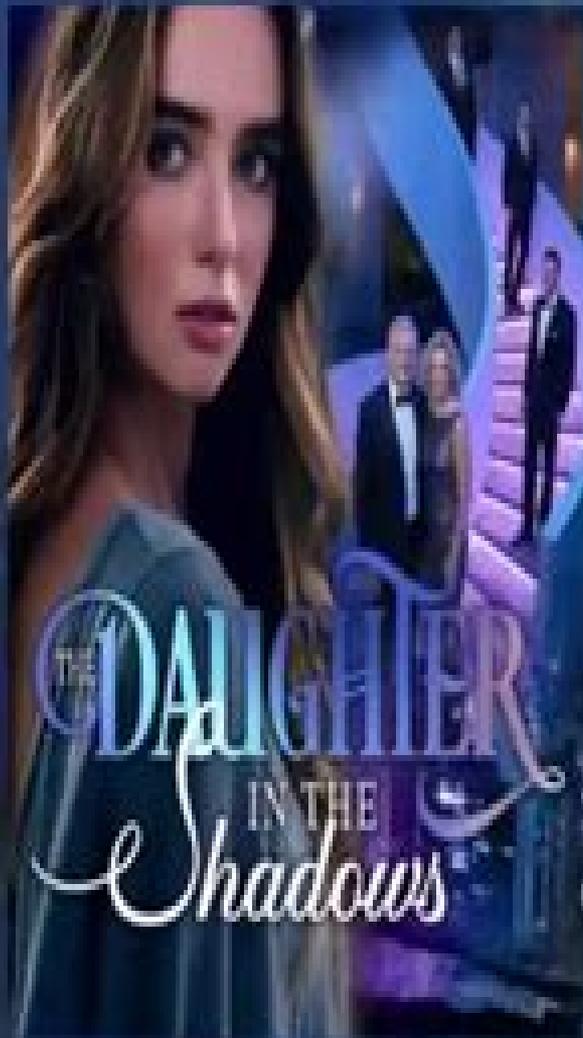


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 $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \frac{5}{2} \text{O}_2 \text{ (g)}$ $\Delta h = 230 \text{ kJ}$
 $2 \text{ H}_2 \text{ (g)} + \text{C}_2\text{H}_6 \text{ (g)} \rightarrow \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} + \frac{1}{2} \text{O}_2 \text{ (g)} \rightarrow \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (g)}$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4 \text{ (g)} + 3 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 2 \text{ H}_2\text{O (l)}$ $\Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (l)}$ $\Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2 \text{ (g)} + \frac{1}{2} \text{O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O (l)}$ $\Delta h = 285 \text{ kJ/mole}$

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