



Chicks And Chickens

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web 33mol O_2 6mol Br_2 11mol O_2 18mol Br_2 identify the limiting reactant s and excess reactant s the limiting reactant is O_2 since it would yield the least amount of product 18 mol Br_2 the excess reactant is $\text{C}_2\text{H}_3\text{Br}_3$ since its complete reaction would have yielded up to 54 mol Br_2 think about your result

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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less mgo mgo than does O_2 0.2398 g mgo vs 25.2 g mgo therefore mg is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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 $3^5 = 43x^2$ $1.6 = 42p^4$ $2p^1 = 7$ $6 = 2a$ $62 = 3a$ $8 = 22x^2$ $23x = 9$ $63m = 6$ $m = 6$ $2m = 10$ $2x = 2$ $2x = 11$ $10 = 3x$ $10x = 1$ $10 = 12$ $3 = 2x$ $1 = 3$ $2x = 3$ $3 = x$

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web step 1 write all logarithmic expressions as a single logarithm with coefficient 1 in this case apply the product rule for logarithms $\log_2 x^2 \log_2 x^3 = 1$ $\log_2 x^2 \times 3 = 1$ step 2 use the definition and rewrite the logarithm in exponential form $\log_2 x^2 \times 3 = 1$
 $3 = 1 \times 2 \times 3 = 2^1$

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