

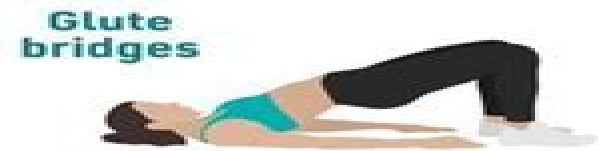


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# Works Proven Strategy For Home Workout Routine Tools Comparison With

**SA Dillow**



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web when the reactants have more potential energy than the products have and the extra energy is released to the surroundings it is called activation energy the height of the hill on a potential energy diagram is a measure of the endothermic the reaction above can be categorized as

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web this is like a one dimensional system whose mechanical energy  $e$  is a constant and whose potential energy with respect to zero energy at zero displacement from the spring  $s$  unstretched length  $x_0$  is  $u x^2 = \frac{1}{2} kx^2$  figure 8 5 2 8 5 2 a a glider between springs on an air track is an example of a horizontal mass spring system

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web 1 if the  $pe$  of the reactants is 430kJ and the  $pe$  of the products is 190kJ what is the  $\Delta h$  2 if the  $pe$  of the reactants is 430kJ and the activation energy is 220 how much total energy will be needed to be added for the reaction to proceed transcribed image text potential energy diagram answer the questions using the graph

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web figure 5 5 1 a potential energy diagram shows the total potential energy of a reacting system as the reaction proceeds a in an endothermic reaction the energy of the products is greater than the energy of the reactants and  $\Delta h$  is positive b in an exothermic reaction the energy of the products is lower than the energy of the

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web pertanyaan serupa iklan perhatikan diagram energi potensial berikut pada diagram di atas energi aktivasi untuk reaksi baliknya ditandai dengan

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web answer see below explanation the rate of a chemical reaction  $a + b \rightarrow c + d$  is given in first approximation by  $v = k_v a^m b^n$  where  $k_v$   $k_e$   $e_a$   $r$  where  $k$  depends by many factors as solvent geometry type of reaction  $e_a$  is the activation energy  $r$  is the universal constant of gases  $t$  is the absolute temperature

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web mar 17 2016 3 what is the activation energy of a reaction and how is this energy related to the activated complex of the reaction  $e_a$  is the minimum amount of energy for a reaction to occur it is the amount of energy required to create an activated complex 4 what happens when a catalyst is used in a reaction a catalyst changes the reaction

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web all questions are answered herein chemistry tutorial 9 02b potential energy diagrams how to draw and label pe diagrams for exothermic and endothermic reactions and what effect adding a catalyst or inhibitor has on the diagram

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