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constant expression for this reaction is  $K_a = \frac{[ce h ce a ce ha label 8]}{[ce h ce a ce ha label 8]}$  where we denote the equilibrium constant  $K$  with a subscript  $a$  for acid

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web equilibrium moles  $hscn = 2.00 \times 10^{-5} \text{ mol}$   $3.00 \times 10^{-6} \text{ mol}$   $1.70 \times 10^{-5} \text{ mol}$   $hscn$  knowing the number of moles of  $fe^{3+}$  and  $hscn$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species  $[fe^{3+}] = 1.70 \times 10^{-5} \text{ mol/l}$   $[fe^{3+}] = 3.850 \times 10^{-4} \text{ m}$

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