

Atomic Structure

Electronic Structure of the atom:

The electronic structure of an atom that determines its chemical properties. Electrons can be regarded as having certain energy levels. The lower levels fill before the higher levels are occupied. The simple theory you learnt at IGCSE, pictures the atom as having shells, the shell of lowest energy being closest to the nucleus. The next shell can hold up to eight electrons. The third shell also holds up to eight electrons.

Bonding is explained in terms of atoms trying to obtain a stable electrons arrangement by:

- (i) transfer from electron rich atoms (metals) to electron deficient atoms (non-metals);
- (ii) sharing electrons between electron deficient atoms. This model on its own is not sophisticated enough for understanding A level work fully and does not explain how the atoms arrange electrons above a number of twenty. The simple bonding theory which comes from the model cannot explain the structure of molecules such as PCl_5 or SF_6 .

Electronic Configuration

- The shells used in the simple view of the atom are divided into subshells.

Shell	Subshells
1	s
2	s p
3	s p d
4	s p d f

- The higher the energy of the shell, the more subshells it contains.

Electrons inside the atom should be thought of as waves rather than particles. The area occupied by an electron wave is called the orbital. An orbital has no edge: the electron density just tails off more and more. However, we can calculate the surface which would contain a given percentage (say, 95%) of the electron density, and use this to gain an idea of the shape of the orbital (below).

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