

# NITROGEN AND ITS COMPOUNDS

Nitrogen is a non metal in group V of the periodic table. It has an atomic number of seven and has an electron arrangement of 2.5. Nitrogen exists in air as a gaseous element consisting of diatomic molecules. About 78% by volume of the atmosphere is composed of nitrogen making it the most important source of nitrogen.

Nitrogen also occurs combined in Compounds such as sodium nitrate, potassium nitrate and proteins

## Nitrogen

### Isolation of Nitrogen from Air

#### Laboratory Isolation

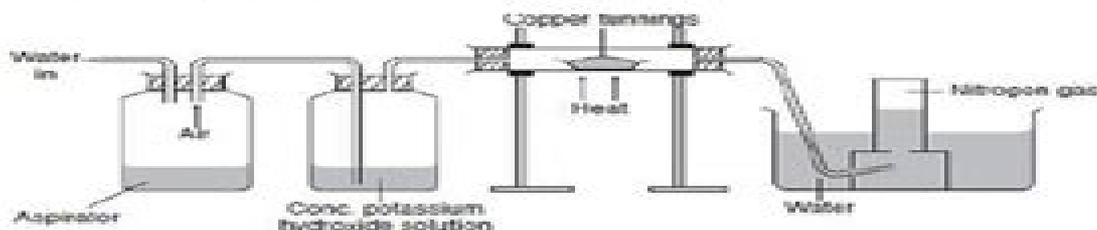
In the laboratory, nitrogen is isolated from air by first passing the air through concentrated potassium hydroxide or sodium hydroxide which **absorbs carbon(IV) oxide**.



The remaining part of air is then passed over heated copper turnings to remove oxygen.



The residual gas is then collected as shown below.



The nitrogen so obtained **contains noble gases as impurities**.

#### Fractional distillation of liquefied air

Nitrogen is obtained on a large scale by fractional distillation of liquid air.

**Dust particles** are first removed from air through **electrostatic precipitation**.

The dust-free air is then passed through **concentrated sodium hydroxide solution to remove carbon(IV) oxide gas**.

The remaining air is passed into a **condensation chamber and cooled to  $-25^\circ\text{C}$** . **Water vapour separates out as ice**. The residual gas is compressed to a pressure of 200 atmospheres.

Repeated expansion and contraction of the air cools it to a liquid at a temperature of  $-200^\circ\text{C}$ . The main constituents of liquid air at this stage are nitrogen and oxygen whose boiling points are  $-196^\circ\text{C}$  and  $-183^\circ\text{C}$  respectively. **Nitrogen has a lower boiling point and therefore distills out first followed by oxygen**.

# Analytical Chemistry Of Nitrogen And Its Compounds

**Clemens Wendtner**



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