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College of Engineering

Department of Civil Engineering

SCI 401: General Chemistry
 Practice Set

1. Consider the following chemical reaction:



What is the change in enthalpy of this reaction?

Solution: Use the values in the Standard Thermodynamic Values and calculate their change in enthalpy using:
 $H = H_{\text{prod}} - H_{\text{react}}$

$\text{HCOH}: H = 228.28 \text{ kJ/mol}$

$\text{HCl}: H = -92.29909 \text{ kJ/mol}$

$\text{H}_2\text{CClCl}: H = 21.96 \text{ kJ/mol}$

The thermodynamic values might differ depending on the source. That's okay. They are almost equal / within the range. As long as you know the process, you're good.

The equation is already balanced so we can proceed with the solution.

$$\Delta H = H_{\text{product}} - H_{\text{reactant}}$$

$$= \left[(21.96 \frac{\text{kJ}}{\text{mol}}) \right] - \left[(228.28 \frac{\text{kJ}}{\text{mol}}) + (-92.29909 \frac{\text{kJ}}{\text{mol}}) \right]$$

$$\Delta H = -114.12096 \frac{\text{kJ}}{\text{mol}}$$

2. Determine ΔH°_f of ethanol ($\text{C}_2\text{H}_5\text{OH}$) from its standard enthalpy of combustion (-1367.4 kJ/mol to produce $\text{CO}_2(\text{g})$ and $\text{H}_2\text{O}(\text{l})$).

Since the reaction is combustion, automatically the substance will react with O_2 gas and produce $\text{CO}_2(\text{g})$ and $\text{H}_2\text{O}(\text{l})$ as products. We need to balance the equation first.



$\text{O}_2(\text{g}): H = 0 \text{ kJ/mol}$
 $\text{CO}_2(\text{g}): H = -393.5052 \text{ kJ/mol}$
 $\text{H}_2\text{O}(\text{l}): H = -285.82996 \text{ kJ/mol}$

$$\Delta H = H_{\text{product}} - H_{\text{reactant}}$$

$$-1367.4 \frac{\text{kJ}}{\text{mol}} = \left[(2)(-393.5052 \frac{\text{kJ}}{\text{mol}}) + (3)(-285.82996 \frac{\text{kJ}}{\text{mol}}) \right] - \left[H_{\text{C}_2\text{H}_5\text{OH}} + 3(0 \frac{\text{kJ}}{\text{mol}}) \right]$$

$$H_{\text{C}_2\text{H}_5\text{OH}} = -277.1003 \frac{\text{kJ}}{\text{mol}}$$

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