
CHEM 121 Lab Report 6 Equilibrium Reactions - Portage Learning

Name:

Date:

Experiment: 6

Title: Equilibrium Reactions

Purpose: the purpose of this lab is to study the equilibrium reactions by doing two experiments including the determination of the equilibrium constant K_c and the study of LeChatelier's Principle

Procedure:

In the small 10 ml flask that was used was filled with acetic acid and then 10 ml of propyl alcohol was added to the flask then it was mixed. Several drops of phenolphthalein indicator is added to the beaker. Then it was titrated with sodium hydroxide. Then 1 ml of it was added to a 25 ml beaker. After titration it was to become a pink color.

Next in the constant experiment we have the mixture of acetic acid and propyl alcohol and sulfuric acid catalyst. It is going to be heated for an hour so the reaction can be encouraged. The acetic acid and the propyl alcohol will react with each other to form propyl acetate + water. Which is the equilibrium. We then measure how much acetic acid is present knowing that some will be used up.

Data/Results/Calculations:

Data Table:

CH₃COOH 10.00ml d=1.05g/ml MW=60

C₃H₇OH 10.00 ml d=0.8 g/ml MW=60

1st titration=17.1ml to check initial CH₃COOH

2nd titration=17.5 ml to account for H₂SO₄

3rd titration= 8.2 ml to measure final CH₃COOH

NaOH M= 0.508 mol/liter

Initial - CH₃COOH = 10.00 ml × 1.05 g/ml / 60 g/mol × 0.020 liter = 8.75 M

Initial - C₃H₇OH = 10.00 ml × 0.80 g/ml / 60 g/mol × 0.020 liter = 6.67 M

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Shirley M. Young



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