

April 14: Chemical Equilibrium

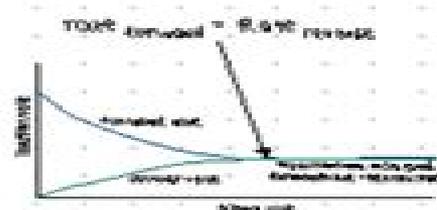
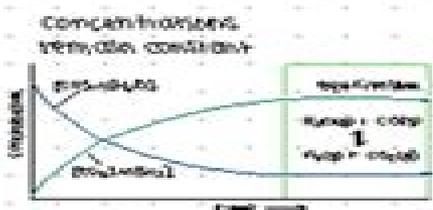
Chemical equilibrium: dynamic process in which concentrations of reactants + products remain constant over time

rate of reaction in forward direction matches rate in reverse direction - forward + reverse rates not zero at equilibrium



dynamic equilibrium: state of chemical reaction in which reactant and product concentrations do not change, but products and reactants are continuously interconverting

Equilibrium



EXAMPLE: The Haber process



1. Initially, only N_2 and H_2 collide and react at some rate
2. as NH_3 forms, there is less N_2 and H_2 therefore less likely to collide... less likely to react
3. equilibrium concentrations are constant, rate of NH_3 dissociation + formation are equal



from experimental data

$$\text{rate}_f = k_f [\text{NO}_2]^2$$

$$\text{rate}_r = k_r [\text{N}_2\text{O}_4]$$

at equilibrium: $k_f [\text{NO}_2]^2 = k_r [\text{N}_2\text{O}_4]$

rearrange: $\frac{k_f}{k_r} = \frac{[\text{N}_2\text{O}_4]}{[\text{NO}_2]^2} = K_{eq}$ (equilibrium constant)

Equilibrium Constant

Equilibrium constant expression:

ratio of equilibrium concentrations or partial pressures of products to reactants each term raised to a power equal to the coefficient of that substance in the balanced chemical equation

Equilibrium constant:

value of the ratio of concentrations (or partial pressures) terms in the equilibrium constant expression at a specific temperature

for general overall reaction



equilibrium constant defined as:

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

where [A], [B], etc are equilibrium concentrations

$K \gg 1$ → equilibrium lies to right; PRODUCTS DOMINATE

$K \ll 1$ → equilibrium lies to left; REACTANTS DOMINATE

$K = 0$ → no reaction

$K = \infty$ → irreversible reaction

LAW OF MASS ACTION

equilibrium constant expression of chemical equilibrium has characteristic value @ given temperature

mass action expression:

equivalent to equilibrium constant expression, but applied to reaction mixtures that may or may not be at equilibrium



Experiment	Starting Concentrations	Equilibrium Concentrations	Calculated value of K_c
1	$[\text{H}_2] = 0.2000 \text{ M}$ $[\text{CO}] = 0.2000 \text{ M}$	$[\text{H}_2] = 0.1875 \text{ M}$ $[\text{CO}] = 0.1875 \text{ M}$ $[\text{CH}_4] = 0.0125 \text{ M}$ $[\text{H}_2\text{O}] = 0.0125 \text{ M}$	$K_c = 0.04$
2	$[\text{H}_2] = 0.2000 \text{ M}$ $[\text{CO}] = 0.2000 \text{ M}$	$[\text{H}_2] = 0.1500 \text{ M}$ $[\text{CO}] = 0.1500 \text{ M}$ $[\text{CH}_4] = 0.0500 \text{ M}$ $[\text{H}_2\text{O}] = 0.0500 \text{ M}$	$K_c = 0.04$
3	$[\text{H}_2] = 0.2000 \text{ M}$ $[\text{CO}] = 0.2000 \text{ M}$	$[\text{H}_2] = 0.1625 \text{ M}$ $[\text{CO}] = 0.1625 \text{ M}$ $[\text{CH}_4] = 0.0375 \text{ M}$ $[\text{H}_2\text{O}] = 0.0375 \text{ M}$	$K_c = 0.04$

equilibrium constant expression: $K_c = \frac{[\text{CH}_4][\text{H}_2\text{O}]}{[\text{H}_2]^3[\text{CO}]}$

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Lei Huang



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Table of Contents Custom Published Chemistry 105

1. Understanding the eBook Custom Published Chemistry 105
 - The Rise of Digital Reading Custom Published Chemistry 105
 - Advantages of eBooks Over Traditional Books
2. Identifying Custom Published Chemistry 105
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Custom Published Chemistry 105
 - User-Friendly Interface
4. Exploring eBook Recommendations from Custom Published Chemistry 105
 - Personalized Recommendations
 - Custom Published Chemistry 105 User Reviews and Ratings
 - Custom Published Chemistry 105 and Bestseller Lists

-
5. Accessing Custom Published Chemistry 105 Free and Paid eBooks
 - Custom Published Chemistry 105 Public Domain eBooks
 - Custom Published Chemistry 105 eBook Subscription Services
 - Custom Published Chemistry 105 Budget-Friendly Options
 6. Navigating Custom Published Chemistry 105 eBook Formats
 - ePub, PDF, MOBI, and More
 - Custom Published Chemistry 105 Compatibility with Devices
 - Custom Published Chemistry 105 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Custom Published Chemistry 105
 - Highlighting and Note-Taking Custom Published Chemistry 105
 - Interactive Elements Custom Published Chemistry 105
 8. Staying Engaged with Custom Published Chemistry 105
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Custom Published Chemistry 105
 9. Balancing eBooks and Physical Books Custom Published Chemistry 105
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Custom Published Chemistry 105
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Custom Published Chemistry 105
 - Setting Reading Goals Custom Published Chemistry 105
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Custom Published Chemistry 105
 - Fact-Checking eBook Content of Custom Published Chemistry 105
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning

-
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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