

CHEM 101 – Recitation Week 2

Ch. 1, 37. Classify each substance as a pure substance or a mixture. If it is a pure substance, classify it as an element or a compound. If it is a mixture, classify it as homogeneous or heterogeneous.

Difference between pure substances and mixtures

Difference between homogeneous and heterogeneous mixtures

- a. sweat
- b. carbon dioxide
- c. aluminum
- d. vegetable soup

Ch. 1, 40. Complete the table.

Comparing pure substances and mixtures

Comparing elements and compounds

Substance	Pure or Mixture	Type
water	pure	compound
coffee		
ice		
carbon		

Ch. 1, 48. An automobile gas tank holds 21 kg of gasoline. When the gasoline burns, 84 kg of oxygen is consumed, and carbon dioxide and water are produced. What is the total combined mass of carbon dioxide and water that is produced?

Law of mass conservation

Law of definite composition

Ch. 1, 52. Upon decomposition, one sample of magnesium fluoride produces 1.65 kg of magnesium and 2.57 kg of fluorine. A second sample produces 1.32 kg of magnesium. How much fluorine (in grams) is formed?

Law of multiple proportions

Dimensional analysis and unit conversions

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Léo Pomar



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