

## Study Guide: Chemical Reactions

### I. Fundamentals

- **What is a chemical reaction?**
  - Process where atoms are rearranged to create new substances with different properties.
  - Involves breaking and forming chemical bonds.
- **Evidence of a chemical reaction:**
  - Color change
  - Formation of a precipitate (a solid formed from liquids)
  - Gas production (bubbles)
  - Temperature change (heat absorbed or released)
  - Light emission
  - Odor change

### II. Chemical Equations

- **Reactants:** Starting materials in a reaction (left side of the arrow).
- **Products:** Substances formed in a reaction (right side of the arrow).
- **Balanced equations:** Number of atoms of each element is equal on both sides of the arrow.
  - Coefficients are used to balance equations.
- **States of matter:** (s) solid, (l) liquid, (g) gas, (aq) aqueous (dissolved in water)

### III. Types of Chemical Reactions

- **Synthesis:** Two or more substances combine to form a single product.
  - $A + B \rightarrow AB$
- **Decomposition:** A single compound breaks down into two or more simpler substances.
  - $AB \rightarrow A + B$
- **Single replacement:** One element replaces another element in a compound.
  - $A + BC \rightarrow AC + B$
- **Double replacement:** The cations and anions of two ionic compounds switch places.
  - $AB + CD \rightarrow AD + CB$
- **Combustion:** A substance reacts with oxygen, often producing heat and light.

### IV. Energy Changes in Chemical Reactions

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**Terry C. Jones**



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