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web solution  $v_1 = 2.50 \text{ l}$   $v_2 = ?$   $t_1 = 273 \text{ K}$   $t_2 = 303 \text{ K}$  we will be applying charles law as pressure is constant here  $v_1 t_1 = v_2 t_2$  or  $v_2 = \frac{v_1 t_1}{t_2} = \frac{2.50 \text{ l} \times 273 \text{ K}}{303 \text{ K}} = 2.25 \text{ l}$  at  $90^\circ \text{C}$  a helium sample has a volume of  $500 \text{ ml}$  determine the temperature at which the volume of the liquid will become  $240 \text{ ml}$