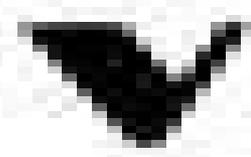


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titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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 $HCl + NaOH \rightarrow NaCl + H_2O$ one mole of HCl would be fully neutralized by one mole of $NaOH$

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction
 $\text{NaOH} + \text{C}_6\text{H}_4\text{COOH} \rightarrow \text{C}_6\text{H}_4\text{COONa} + \text{H}_2\text{O}$ the point of equivalence is the point of titration when the number of moles of naoh added is equal to the number of moles of khp in the

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