

# Importance of Custom Clearance Process in International Trade



# About Customs

**SJ Ball**



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web 1 consider the reaction  $i_2O_5 + 5 CO \rightarrow 5 CO_2 + I_2$  80.0 grams of iodine pentoxide reacts with 28.0 grams of carbon monoxide determine the mass of iodine which could be produced b if in the above situation only 0.160 moles of iodine was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key  $4 KO_2 + 2 H_2O \rightarrow 4 KOH + O_2$  g if a reaction vessel contains 0 mol  $KO_2$  and 0 mol  $H_2O$  what is the limiting reactant

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theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$   
 $8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$   
 problems 6 & 8 are similar to problem 5

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 4.73% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  
 $\frac{4.73 \text{ g}}{16.0 \text{ g}} = 0.295 \text{ mole } \text{O}_2$   
 $\frac{21.3 \text{ g}}{32.0 \text{ g}} = 0.666 \text{ mole } \text{C}_3\text{H}_6$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained / calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$   
 4 Al + 3  $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$   
 b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 302.3 \text{ g } \text{Al}_2\text{O}_3$   
 $26.98 \text{ g Al} \times \frac{1 \text{ mol Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 26.98 \text{ g } \text{Al}_2\text{O}_3$   
 c determine the percent yield  
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$   
 theoretical 302.3 g

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighed by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

c is the limiting reagent 4 33 g of h<sub>2</sub> are left over 13

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web the percent yield is calculated as follows text percent yield  $\frac{\text{text actual yield}}{\text{text theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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