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# Cutting Stone

**SA Dillow**



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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction time the total pressure  $p_a = x_a p_t$  mole fraction  $x_a = \frac{\text{moles } a}{\text{total moles}}$  mole fraction of substance

a is often denoted by  $x_a$  and

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web 1 in the given mixture the sum of all the mole fractions is equal to one  $x_a + x_b + \dots = 1$  the multiplication of the mole fraction

by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction

mole fraction is not dependent on the temperature

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web nov 6 2023  $x_a$  mol a mol a mol b and  $x_b$  mol b mol a mol b note in the given mixture the sum of all the mole fractions

present is equal to one  $x_a + x_b + \dots = 1$  when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is

a unitless and dimensionless expression image will be uploaded soon

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ways to represent the concentration of the solution it is defined as the number of molecules of component a divided by the

total number of molecules of all the components in the solution

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 - 30 = 70 g benzene's molar

mass  $C_6H_6 = 78 \text{ g mol}^{-1}$  hence  $C_6H_6$ 's number of moles  $= \frac{30}{78} = 0.3846 \text{ mol}$  molar mass of carbon tet  $CCl_4 =$

$154 \text{ g mol}^{-1}$

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web jan 30 2020 ans the molality of solution is  $1 \text{ mol kg}^{-1}$  or  $1 \text{ m}$  example 02 11.11 g of urea  $NH_2CONH_2$  was dissolved in 100

g of water calculate the molarity and molality of the solution given n 14 h 1 c 12 o 16 given mass of solute urea 11.11 g mass

of solvent water 100 g 0.1 kg to find molarity of solution solution

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web solution moles of  $CH_3OH = \frac{5.532}{32} = 0.173 \text{ mole}$  moles of  $H_2O = \frac{40.18}{18} = 2.23 \text{ moles}$  therefore according to the equation mole fraction

of  $CH_3OH = \frac{0.173}{0.173 + 2.23} = 0.073$  to solve more examples on mole fraction formula and practice more

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moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g kbr

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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xylene  $C_8H_{10}$  has a density of  $0.864 \text{ g/ml}$  a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if  $19.0 \text{ ml}$  of acetone was dissolved in  $98.4 \text{ ml}$  of xylene 7.1

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web since the mole fraction of the  $O_2$  is  $0.050$  it contributes  $0.050$  or if you will  $5.0$  of the total pressure thus total pressure is  $0.2105042 \text{ atm}$  problem 9 a sample of  $1.43 \text{ g}$  of helium and an unweighed quantity of  $O_2$  are mixed in a flask at room temperature