

Ideal Gas Law

The diagram shows the equation $PV = Nk_B T$ with the following labels and arrows:

- Pressure** (green text) with an arrow pointing to P .
- Volume** (blue text) with an arrow pointing to V .
- Number of particles** (black text) with an arrow pointing to N .
- Boltzman's constant** (black text) with an arrow pointing to k_B .
- Temperature** (orange text) with an arrow pointing to T .

$$k_B = 1.38 \times 10^{-23} \text{ m}^2 \text{ kg s}^{-2} \text{ K}^{-1}$$

In words: The product of **pressure** and **volume** is equal to the **number of particles** multiplied by the **Boltzman's constant** and **temperature**.

$$Nk_B = nR, \text{ } n = \# \text{ of moles and } R \text{ is the Gas constant}$$

3d Flow Of Ideal Gases Around Smooth Bod

S Ben Porath



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Scientific and Technical Aerospace Reports, 1987 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database

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