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web a at atmospheric pressure helium can exist in all three phases as well as a supercritical fluid phase near absolute zero choice b solid and gaseous helium never exist in equilibrium with each other at any temperature or pressure b solid and gaseous helium never exist in equilibrium with each other at any temperature or pressure choice c

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web stant is the foundation of thermochemistry it allows us to describe heat flow in terms of the change in a state function enthalpy change is path independent note with regard to choice b that enthalpy change equals internal energy change only if the volume of the system remains constant and thus no thermodynamic work is performed 5 b

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web 1 answer a explanation the value of ΔH is negative therefore the reaction is exothermic or heat releasing 2 answer a 1 explanation the entire reaction was multiplied by 2 therefore ΔH must also be halved 3 answer c explanation option c is incorrect because 566 kj mol is the ΔH per mole of O_2 and not for 2 moles

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy e of a system in any process is equal to the heat q added to the system plus the work w done on the system by its surroundings e

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