

1. Draw Lewis structures, predict the molecular geometry, and give the bond angles for each of the following compounds:

a. Boron trifluoride,  $\text{BF}_3$

b. carbonate ion,  $\text{CO}_3^{2-}$

c. Sulfate ion,  $\text{SO}_4^{2-}$

d. Methane,  $\text{CH}_4$

| 7

2.7

e. Oxygen difluoride,  $\text{OF}_2$

f. Nitrogen trifluoride,  $\text{NF}_3$

g. Hydrogen cyanide,  $\text{HCN}$

2. For each of the Lewis structures you drew in question 1, indicate the hybridization of the central atom.

a. Boron trifluoride,  $\text{BF}_3$

b. carbonate ion,  $\text{CO}_3^{2-}$

c. Sulfate ion,  $\text{SO}_4^{2-}$

d. Methane,  $\text{CH}_4$

e. Oxygen difluoride,  $\text{OF}_2$

f. Nitrogen trifluoride,  $\text{NF}_3$

g. Hydrogen cyanide,  $\text{HCN}$

# Boronnitrogen Compounds

**Clifford Lane**



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