

Department of Chemical Engineering
Series Test 1- CH 010 502 Chemical Reaction Engineering

Time: 1 hr

Maximum Marks:30

Part 1

- 1) At 500 K the rate of a bimolecular reaction is ten times the rate at 400 K. Find the activation energy of this reaction using Collision theory.(2 marks)
- 2) For an exothermic reversible reaction which one of the following correctly describes the dependence of equilibrium constant (K) with temperature (T) and pressure (P) ?
 - a) K is independent of T and P
 - b) K increases with increase in T and P
 - c) K increases with T and decreases with P
 - d) K decreases with an increase in T and is independent of P (1 marks)
- 3) For which reaction order, the half-life of the reactant is half of the full lifetime (time for 100% conversion) of the reactant? (A) Zero order (B) Half order (C) First order (D) Second order (1 marks)
- 4) For the chemical reaction $P \rightarrow Q$, it is found that the rate of reaction doubles as the concentration of 'P' is doubled. If the reaction rate is proportional to C_p^n , then what is the value of 'n' for this chemical reaction ? (A) 1 (B) 2 (C) 3 (D) 0 (1 marks)

Part 2

(3×5 = 15 marks)

- 1) The concentrations of a compound undergoing chemical change were 5.72, 3.23 and 1.96 at the times 0, 20 and 50 min from the commencement of reaction .Ascertain the order of the reaction.
- 2) Discuss various intermediates used in non-elementary chemical reactions
- 3) Briefly explain temperature dependence from Arrhenius theory

Part 3

(10 marks)

Answer any one of the following

- 1) The following data are obtained at 0°C in a constant volume batch reactor using pure gaseous A:

Time in min	0	2	4	6	8	10	12	14	∞
pA in mm Hg	760	60	475	390	320	275	240	215	150

The reaction is $A \rightarrow P$ with $E_a = 2.5$ K .Find the rate equation for this reaction.

OR

- 2) a) Explain Psuedo-steady state hypothesis (PSSH)(3 marks)

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Table of Contents Chemical Engineering Sample Exam

1. Understanding the eBook Chemical Engineering Sample Exam
 - The Rise of Digital Reading Chemical Engineering Sample Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Engineering Sample Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Chemical Engineering Sample Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Engineering Sample Exam
 - Personalized Recommendations
 - Chemical Engineering Sample Exam User Reviews and Ratings

- Chemical Engineering Sample Exam and Bestseller Lists
- 5. Accessing Chemical Engineering Sample Exam Free and Paid eBooks
 - Chemical Engineering Sample Exam Public Domain eBooks
 - Chemical Engineering Sample Exam eBook Subscription Services
 - Chemical Engineering Sample Exam Budget-Friendly Options
- 6. Navigating Chemical Engineering Sample Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Engineering Sample Exam Compatibility with Devices
 - Chemical Engineering Sample Exam Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Engineering Sample Exam
 - Highlighting and Note-Taking Chemical Engineering Sample Exam
 - Interactive Elements Chemical Engineering Sample Exam
- 8. Staying Engaged with Chemical Engineering Sample Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Engineering Sample Exam
- 9. Balancing eBooks and Physical Books Chemical Engineering Sample Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Engineering Sample Exam
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Engineering Sample Exam
 - Setting Reading Goals Chemical Engineering Sample Exam
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Engineering Sample Exam
 - Fact-Checking eBook Content of Chemical Engineering Sample Exam
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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