

## Practice Problems: Quantum Theory

CHEM 1A

1. Hydrogen has a red emission line at 656.3 nm, what is the energy and frequency of a photon of this light? Note: Planck's constant =  $6.626 \times 10^{-34}$  J·s, speed of light =  $2.998 \times 10^8$  m/s

$$E_{\text{photon}} = h\nu = \frac{hc}{\lambda}$$

$$E_{\text{photon}} = \frac{(6.626 \times 10^{-34} \text{ J}\cdot\text{s}) (2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left( \frac{10^{-9} \text{ m}}{1 \text{ nm}} \right)}$$

$$E_{\text{photon}} = \underline{3.02677861 \times 10^{-19} \text{ J}}$$

$$\nu = \frac{c}{\lambda}$$

$$\nu = \frac{(2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left( \frac{10^{-9} \text{ m}}{1 \text{ nm}} \right)}$$

$$\nu = \underline{4.568032912 \times 10^{14} \text{ s}^{-1}}$$

Answer (E): 3.027 x 10<sup>-19</sup> JAnswer (ν): 4.568 x 10<sup>14</sup> s<sup>-1</sup> or Hz

2. What is the de Broglie wavelength for an electron moving at  $2.42 \times 10^6$  m/s? Note:  $m_e = 9.109 \times 10^{-31}$  kg

$$\lambda = \left( \frac{h}{mv} \right)$$

$$\lambda = \left( \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{(9.109 \times 10^{-31} \text{ kg}) (2.42 \times 10^6 \text{ m/s})} \right) = \underline{3.00583657 \times 10^{-10} \text{ m}}$$

Answer: 3.01 x 10<sup>-10</sup> m ← many orders of magnitude larger than the diameter of an electron

3. Using the Bohr equation, calculate is the change in energy when an electron transitions from  $n = 3$  to  $n = 2$  in a hydrogen atom.

$$\Delta E = -2.18 \times 10^{-18} \text{ J} \left( \frac{Z^2}{n_{\text{final}}^2} - \frac{Z^2}{n_{\text{initial}}^2} \right)$$

$$\Delta E = -2.18 \times 10^{-18} \text{ J} \left( \frac{1^2}{2^2} - \frac{1^2}{3^2} \right)$$

$$\Delta E = \underline{-3.02777778 \times 10^{-19} \text{ J}}$$

Answer: -3.03 x 10<sup>-19</sup> J

# Chemistry Theory And Problems

**G Thomas**



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