



$$15.0 \text{ g HI} \times \frac{1 \text{ mol HI}}{128 \text{ g HI}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol HI}} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2}{1 \text{ mol H}_2} \\ = 3.53 \times 10^{22} \text{ molecules H}_2$$



$$12.0 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \times \frac{2 \text{ mol Ca}}{1 \text{ mol O}_2} \times \frac{40.1 \text{ g Ca}}{1 \text{ mol Ca}} = 30.1 \text{ g Ca}$$



$$8.0 \text{ g AgCl} \times \frac{1 \text{ mol AgCl}}{143 \text{ g AgCl}} \times \frac{1 \text{ mol NaCl}}{1 \text{ mol AgCl}} \times \frac{58.5 \text{ g NaCl}}{1 \text{ mol NaCl}} \times \\ \frac{100 \text{ g NaCl soln}}{9.0 \text{ g NaCl}} = 81.8 \text{ g NaCl soln}$$



$$0.320 \text{ mol O}_2 \times \frac{2 \text{ mol H}_2\text{O}_2}{1 \text{ mol O}_2} \times \frac{34.0 \text{ g H}_2\text{O}_2}{1 \text{ mol H}_2\text{O}_2} = 21.8 \text{ g H}_2\text{O}_2$$

$$1.35 \text{ mol Cl}_2 \times \frac{2 \text{ mol Cl}_2\text{O}_3}{2 \text{ mol Cl}_2} \times \frac{119 \text{ g Cl}_2\text{O}_3}{1 \text{ mol Cl}_2\text{O}_3} = 161 \text{ g Cl}_2\text{O}_3$$

5.

$$16.0 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \times \frac{2 \text{ mol Cl}_2\text{O}_3}{3 \text{ mol O}_2} \times \frac{119 \text{ g Cl}_2\text{O}_3}{1 \text{ mol Cl}_2\text{O}_3} = 39.7 \text{ g Cl}_2\text{O}_3$$

# Chemistry Problems And How To Solve Them

**Wolfgang Guggemos**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red-to-white gradient semi-circle overlapping the right side of the bar.

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