

## OXIDATION-REDUCTION REACTIONS

Many parallels can be drawn between acid-base and oxidation-reduction (redox) reactions. Instead of proton interchange and an acid-base conjugate pair, in redox reactions we have electron interchange and a redox conjugate pair



where ox is an electron acceptor or oxidizing agent and red an electron donor or reducing agent. The electron pressure exerted by a conjugate pair is expressed quantitatively in terms of the electrode potential (Table 12-1). Acid and base strengths are leveled (Section 4-3) or limited by the solvent; similarly, oxidants and reductants are leveled by the solvent. In water, strong reductants are leveled in strength (at equilibrium) to that of hydrogen, and oxidants to that of oxygen. Fortunately, for many reagents redox equilibrium with water is attained so slowly that the restriction does not apply in practice.

For acid-base reactions, which are fast, tables of ionization constants are directly applicable. In contrast, for redox reactions, many of which are slow, a table of electrode potentials can be no more than a guide to equilibrium conditions as it tells nothing of rates of reactions or mechanisms. For example, a table of electrode potentials leads to the expectation of a quantitative reaction between Ce(IV) and

# Chemical Analysis An Advanced Text And Reference

**SJ Ball**



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