

# Some Basic Principles of Chemistry

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1. Suppose the elements  $X$  and  $Y$  combine to form two compounds  $XY_2$  and  $X_3Y_2$ . When 0.1 mole of  $XY_2$  weighs 10 g and 0.05 mole of  $X_3Y_2$  weighs 9 g, the atomic weights of  $X$  and  $Y$  are [NEET 2016 Phase II]  
(a) 40, 30 (b) 60, 40 (c) 20, 30 (d) 30, 20
2. If Avogadro number  $N_A$  is changed from  $6.022 \times 10^{23} \text{ mol}^{-1}$  to  $6.022 \times 10^{20} \text{ mol}^{-1}$  this would change [CBSE AIPMT 2015]  
(a) the definition of mass in units of grams  
(b) the mass of one mole of carbon  
(c) the ratio of chemical species to each other in a balanced equation  
(d) the ratio of elements to each other in a compound
3. 20.0 g of a magnesium carbonate sample decomposes on heating to give carbon dioxide and 8.0 g magnesium oxide. What will be the percentage purity of magnesium carbonate in the sample? (Atomic weight of Mg = 24) [CBSE AIPMT 2015]  
(a) 75 (b) 96 (c) 60 (d) 84
4. What is the mass of precipitate formed when 50 mL of 16.9% solution of  $\text{AgNO}_3$  is mixed with 50 mL of 5.8%  $\text{NaCl}$  solution? (Ag = 107.8, N = 14, O = 16, Na = 23, Cl = 35.5) [CBSE AIPMT 2015]  
(a) 28 g (b) 3.5 g (c) 7 g (d) 14 g
5. The number of water molecules is maximum in [CBSE AIPMT 2015]  
(a) 18 molecules water (b) 1.8 g of water  
(c) 18 g of water (d) 18 moles of water
6. When 22.4 L of  $\text{H}_2(\text{g})$  is mixed with 11.2 L of  $\text{Cl}_2(\text{g})$ , each at STP, the moles of  $\text{HCl}(\text{g})$  formed is equal to [CBSE AIPMT 2014]  
(a) 1 mole of  $\text{HCl}(\text{g})$  (b) 2 moles of  $\text{HCl}(\text{g})$   
(c) 0.5 mole of  $\text{HCl}(\text{g})$  (d) 1.5 moles of  $\text{HCl}(\text{g})$
7. 1.0 g of magnesium is burnt with 0.56 g of oxygen in a closed vessel. Which reactant is left in excess and how much? [CBSE AIPMT 2014]  
(At. weight of Mg = 24, O = 16)  
(a) Mg, 0.16 g (b)  $\text{O}_2$ , 0.16 g  
(c) Mg, 0.44 g (d)  $\text{O}_2$ , 0.28 g
8. How many grams of concentrated nitric acid solution should be used to prepare 250 mL of 2.0 M  $\text{HNO}_3$ ? The concentrated acid is 70%  $\text{HNO}_3$ . [NEET 2013]  
(a) 45.0 g conc.  $\text{HNO}_3$  (b) 90.0 g conc.  $\text{HNO}_3$   
(c) 70.0 g conc.  $\text{HNO}_3$  (d) 54.0 g conc.  $\text{HNO}_3$
9.  $6.02 \times 10^{20}$  molecules of urea are present in 100 mL of its solution. The concentration of solution is [NEET 2013]  
(a) 0.02 M (b) 0.01 M  
(c) 0.001 M (d) 0.1 M

# Chem Principles Sm

**Camilla Rothe**



**Chem Principles Sm:**

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Chem Principles Sm**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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