

Types of Isomers:

Constitutional Isomers- have the same molecular (empirical) formulas but have different connectivity's- the same parts but arranged in different ways. Ex. 1-hexene, 2-hexene.

Conformational Isomers- are stereoisomers that can be converted into one another by rotation around a single bond.

Configurational Isomers- Are stereoisomers that cannot be converted into one another by rotation around a single bond.

Geometric isomers: are molecules that are locked into their spatial positions with respect to one another due to a double bond or a ring structure. **Optical isomers:** are molecules that differ three-dimensionally by the placement of substituents around one or more atoms in a molecule. (some into the page some out of the page. **Stereoisomers-** have the same

connectivity but a different arrangement of their atoms in space. Ex. E/Z isomers, R/S isomers.

Enantiomers- stereoisomers that are non-superimposable mirror images

Non-superimposable mirror image- super-imposable molecules are considered to be the same molecule as if reflected in a mirror both give the same molecule.

Naming structures: Meth, Eth, Prop, But, Pent, Hex, Hept, Oct, Non, Dec

Leaving groups- three molecules are recognized as great leaving groups: Cl<Br<I

Rules: CIP- to name a specific stereoisomer of a molecule. Find the chiral carbon- the carbon with attached to four

differing groups, arrange in order of priority, highest priority being the atom with the highest atomic number. **Saytzeff rule-** the major product is the most substituted alkene (least number of hydrogens on double bonded carbons, that is the alkene with the least number of hydrogens directly attached to the carbons of the C=C.

Markownikoffs rule¹- addition of an unsymmetrical reagent to an unsymmetrical alkene, gives as the major product the compound in which the electropositive part of the reagent (usually H⁺) has bonded to the carbon of the C=C that is directly bonded to the greater number of hydrogen atoms.

Polar/Non-Polar- when there are no polar bonds in a molecule, there is no permanent charge difference between one part of the molecule and another and the molecule is considered to be non-polar.

Ortho, Para, Meta (deactivating, activating)- activating groups increase the rate of reaction, relative to H. deactivating groups decrease the rate of reaction in relation to H.

E/Z isomers- Record the groups highest priority and lowest priority on each side if both the highest priority groups sit on the same side then the compound is a Z compound, opposite sides mean a E compound.

R/S isomers- after using the CIP rule and determining the chiral carbon/stereogenic center and groups from highest to lowest priority, put the 4 groups sitting around the stereogenic center and determine whether the groups priority is going from 1-4 clockwise or anticlockwise. Clockwise results in a R configuration, Anticlockwise results in a S configuration.

Curly arrows- Curly arrows always flow from an electron-rich species to an electron-poor species. I.e. from the nucleophile to electrophile. They get a start from lone pairs or bonds and end up either between a pair of atoms or an atom. **Average rate vs instantaneous rate-** the average rate is the change in rate over a time interval. However instantaneous rate is the rate at a specific point in time during the course of the reaction.

| SUBSTRATE | REAGENT TYPE | | |
|-----------|---|---|---|
| | Weaker Nucleophile (low density) | Stronger Nucleophile (low density) | Stronger Nucleophile (high density) |
| Primary | S _N 2 | S _N 2 | S _N 2 |
| Secondary | S _N 1 or S _N 2 | S _N 2 | S _N 2 & E2 |
| Tertiary | S _N 1 & E1 | S _N 1 & E1 | E2 |
| | Typical Nucleophile OH ⁻ , O ²⁻ | Stronger Nucleophile Cl ⁻ , Br ⁻ , I ⁻ , CN ⁻ , CH ₃ ⁻ | Strongest Nucleophile H ⁻ , CH ₃ ⁻ , C(CH ₃) ₃ ⁻ , C(CH ₃) ₂ ⁻ |

Chair conformers: is a cyclohexane somewhat resemble a chair. The

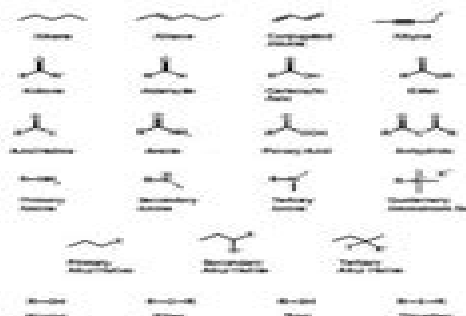
positioned with the bonds pointing up attached to the chair conformer are considered to be on the same side therefore are cis to each other. If the groups up and down, then they are trans.

Axial- means the groups on the chair conformer that are sitting on the **Equatorial-** are groups that are at conformer.

Stable Chair conformers: the more more stable the chair conformer.

Pi bonds: are types of covalent bonds they are generally weaker than sigma in double bonds and twice in triple

Major Organic Chemistry Functional Groups



ring that has been altered in shape to functional groups on the chair are or down. If the majority of the groups pointing down, then they are



corners angles of in the middle of the

groups in the equatorial position the

that are two p-orbitals overlapping, bonds and are found appearing once bonds.

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