

INORGANIC QUALITATIVE ANALYSIS

	EXPERIMENT	OBSERVATION	INFERENCE
1.	<p><u>PRELIMINARY ANALYSIS</u></p> <p>Noted Colour and appearance of salt</p>	<p>a) Colourless crystals</p> <p>b) Colourless amorphous</p>	<p>Absence of Cu^{2+}, Fe^{2+}, Mn^{2+}, etc.</p> <p>Maybe carbonates</p>
2.	<p>Noted the solubility in water</p>	<p>a) Soluble</p> <p>b) Insoluble</p>	<p>May be halides, Nitrates, sulphates</p> <p>May be carbonates</p>
3.	<p>Flame test</p> <p>A little of salt is made into a paste with conc. HCl. A little of this paste is taken on a glass rod and shown to the non-luminous Bunsen flame.</p>	<p>a) Pale green flame</p> <p>b) Brick red flame</p> <p>c) Crimson red flame</p>	<p>Presence of Ba^{2+}</p> <p>Presence of Ca^{2+}</p> <p>Presence of Sr^{2+}</p>
4.	<p>Ash Test</p> <p>A little of the salt is mixed with a few drops of conc. HNO_3 and cobalt nitrate solution. A multi-folded filter paper is soaked in this and</p>	<p>a) A blue tinted ash</p> <p>b) A green tinted ash</p> <p>c) A pink tinted ash</p>	<p>Presence of Al^{3+}</p> <p>Presence of Zn^{2+}</p> <p>Presence of Mg^{2+}</p>

Chemistry The Central Science Qualitative Inorganic Analysis

G Orfield



Chemistry The Central Science Qualitative Inorganic Analysis:

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