

CHEMISTRY

Year Group	10	Examination Length	1 hr 15 mins	Total Marks	100
How you will be assessed	Mixture of Multiple choice, Short answers and Long answer questions.				
Exam Details	AQA Code Code 8462				
Equipment List	Black Pen, Sharp Pencil, Eraser, Calculator, Ruler				

What you will be assessed on:

1. Atomic structure and the periodic table

- Atoms, Elements, molecule, Compounds (Define each term and give examples using formulae and diagrams)
- Naming compounds (example FeS and FeSO₄, be able to identify the number of atoms in a compound)
- Separating Mixtures - (know the physical property used to separate the mixture, Describe the techniques using labelled diagrams and key words - distillation, filtration, crystallisation)
- Relative mass charge and location of subatomic particles (Describe the structure of the atom and relate this to the periodic table using C as an example)
- Isotopes and Relative atomic mass calculation (define isotope and know the formula to calculate the RAM)
- Electronic structure (be able to draw a labelled diagram to represent an atom example neon) (How do you deduce the electronic configuration of an atom)
- Development of the model of the atom (Dalton, Thompson, Rutherford, Bohr, Chadwick)
- Draw the model of the atom related to each scientist and describe the model
- Give details of the Rutherford experiment that led to the conclusion of the new model of the atom
- What was the significance of Chadwick's model
- The Periodic Table (describe Mendeleev's work and how is this different to the modern day periodic table)
Annotate the blank periodic table to show your understanding of periods groups, position of alkali metals, halogen, transition metals and noble gases.
Why are noble gases unreactive?
What does the period number tell you?
What does the group number tell you?
In what order are the elements in the periodic table arranged
- Group 1 (know the names of the first three alkali metals. Explain why alkali metals are stored in oil. Write a word and balanced equation for this.
- Write a word and balanced equation when alkali metals react with water. State the observations. Compare these observations as you go down the group.
- State and explain the trend in reactivity.
- Write a word and balanced equation when alkali metals react with halogens. State the colour of the compound)
- Transition metals (typical properties) (compare the physical and chemical properties with alkali metals)
- What is a Physical Property - Idea of a feature of a substance that can be observed or measured ✓
- Group 7 (describe the state and colour of first four halogens. State how and explain the trend in reactivity.
- What is a displacement reaction give an example with the observations. Give a balanced equation)

Chemistry Year 1

University of Wisconsin



Chemistry Year 1:

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Table of Contents Chemistry Year 1

1. Understanding the eBook Chemistry Year 1
 - The Rise of Digital Reading Chemistry Year 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Year 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Year 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Year 1

- Personalized Recommendations
 - Chemistry Year 1 User Reviews and Ratings
 - Chemistry Year 1 and Bestseller Lists
5. Accessing Chemistry Year 1 Free and Paid eBooks
 - Chemistry Year 1 Public Domain eBooks
 - Chemistry Year 1 eBook Subscription Services
 - Chemistry Year 1 Budget-Friendly Options
 6. Navigating Chemistry Year 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Year 1 Compatibility with Devices
 - Chemistry Year 1 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Year 1
 - Highlighting and Note-Taking Chemistry Year 1
 - Interactive Elements Chemistry Year 1
 8. Staying Engaged with Chemistry Year 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Year 1
 9. Balancing eBooks and Physical Books Chemistry Year 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Year 1
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Year 1
 - Setting Reading Goals Chemistry Year 1
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Year 1

- Fact-Checking eBook Content of Chemistry Year 1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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